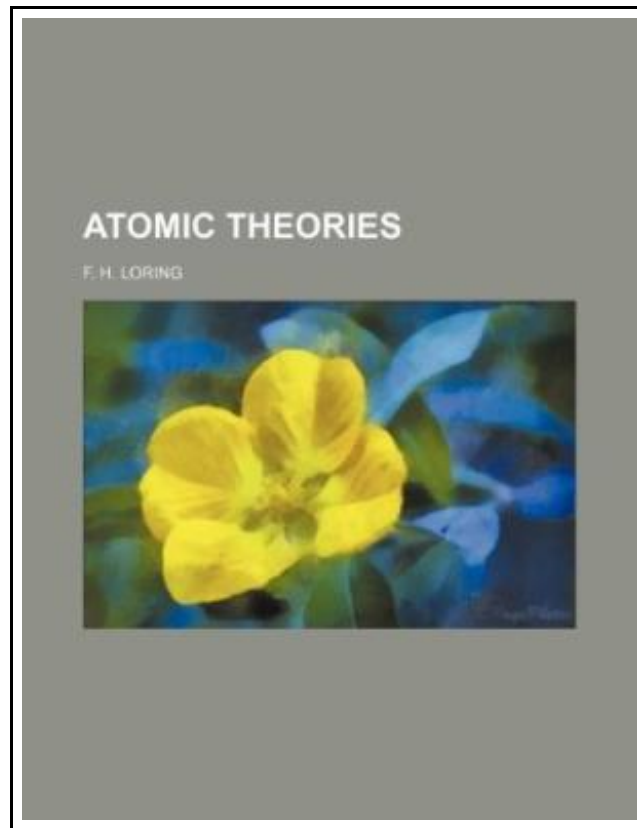


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Reviews

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ATOMIC THEORIES (PAPERBACK)



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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1921 Excerpt: .up enough electrons to complete its octet. 11. The octets may hold one, two or sometimes three pairs of electrons in common. One octet may share 1, 2, 3 or 4 pairs of its electrons with 1, 2, 3 or 4 other octets. One or more pairs of electrons in an octet may be shared by the corresponding number of hydrogen nuclei. No single electron can be shared by more than two octets. The principle of sharing electrons in a rather definite way was proposed by J. Stark.f who suggested that when chemical combination occurs between two atoms the attraction takes place through an intermediate electron which forms a bond between the atoms, as illustrated by Fig. 26: this being the simplest form of chemical union. Atoms also have positive patches (+), aa indicated by Fig. 27. These patches are able to bind other atoms by lines of force extending to the electrons of other atoms, as shown by Fig. 28. In this case the number of lines per electron is fewer. These figures are similar to those given by Campbell (loc. cit.). It will be seen that the present theory involves electrons going over entirely from one atom to another in certain cases, bui in others the electrons are shared in much the same way as suggested by Stark. Sir Wm. Ramsay s J theory is also of interest in this See Cohen s Organic Chemistry (1918), I, pp. 60, 82, 85 and 90; also...



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